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The reaction is obviously acid catalyzed (since H^+ affects the
order but does not appear as a reactant) This mechanism in the
steady state approximation gives the correct rate law: $ketone +$
 $H^+ \rightarrow H^+ \cdot ketone$ (slow) $H^+ \cdot ketone + I_2 \rightarrow iodo\text{-}ketone + 2H^+$
 I^- (fast) a) The overall order is 2.

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 $0.001 = - 0.062 \text{ V}$ b) Reversible electrical work is free energy change; molar free energy change is the electrochemical potential difference across the membrane.

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