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Thiosulfate

Saturated Solution

# **Sodium Thiosulfate Saturated Solution**

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**Sodium Thiosulfate**

**Saturated Solution**

Sodium thiosulfate (sodium thiosulphate) is an inorganic compound with the formula  $\text{Na}_2\text{S}_2\text{O}_3 \cdot x\text{H}_2\text{O}$ . Typically it is available as the white or colorless pentahydrate,  $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$ . The solid is an efflorescent (loses water readily) crystalline substance that dissolves well in water.

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thiosulfate is used in gold mining, water treatment, analytical chemistry, the development ...

## **Sodium thiosulfate - Wikipedia**

Get Free Sodium Thiosulfate Saturated Solution Sodium Thiosulfate Saturated Solution Sodium thiosulfate is an inorganic compound with the formula  $\text{Na}_2\text{S}_2\text{O}_3 \cdot x\text{H}_2\text{O}$ .

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Typically it is available as the white or colorless pentahydrate,  $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$ . The solid is an efflorescent crystalline substance that dissolves well in water.

### **Sodium Thiosulfate Saturated Solution - e13 Components**

Sodium thiosulfate |  
 $\text{Na}_2\text{S}_2\text{O}_3$  or  $\text{Na}_2\text{O}_3\text{S}_2$  |  
CID 24477 - structure,  
chemical names,  
physical and chemical



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Thiosulfate

properties,

classification, patents,

literature, biological ...

## **Sodium thiosulfate | Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub> - PubChem**

Sodium thiosulfate

solution c(Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub> · 5

H<sub>2</sub>O) = 0.1 mol/l (0.1

N) Titripur® Reag. Ph

Eur, Reag. USP;

Synonym: Antichlor;

find Supelco-1.09147

MSDS, related peer-

reviewed papers,

technical documents,

similar products &

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more at Sigma-Aldrich.

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**Sodium thiosulfate  
solution c(Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub> 5  
H<sub>2</sub>O) = 0.1 mol/l ...**

Sodium thiosulfate  
solution for 1000 ml,  
c(Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub>) = 0.1 mol/l  
(0.1 N) Titrisol® - Find  
MSDS or SDS, a COA,  
data sheets and more  
information.

**Sodium thiosulfate  
solution | 109950**

According to the  
method, under

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conditions that pH is 11 and temperature is 100DEG C, 21% saturated sodium sulfite solution and sulphur powder are mixed according to the stoichiometric ratio of a...

**CN103922292A -  
Method for  
preparing sodium  
thiosulfate ...**

Suppose you are presented with a clear solution of sodium

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thiosulfate,

$\text{Na}_2\text{S}_2\text{O}_3$

... A solution is

said to be saturated

when it cannot dissolve

any extra solute in it.

**Suppose you are presented with a clear solution of sodium ...**

A solution that is in

equilibrium with

undissolved solute is

called a saturated

solution (there is not

net change in the

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amount of solute that dissolves). Under ...

Sodium thiosulfate: In a clean test tube, place 4 mL of distilled water and 15 grams of sodium thiosulfate, ...

## **Solutions and Solubility**

SODIUM THIOSULFATE SOLUTION. contact cannot be tolerated. Symptoms include coughing, hoarseness, sneezing,tearing, and breathingdifficulty.

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However, workers who cannot escape high accidental exposure may suffer severe pulmonary damage which can be fatal.

Contact with powdered potassium, sodium metals, alkali, and oxidizing agents produce violent reactions.

## **Safety Data Sheet**

**SODIUM**

**THIOSULFATE**

**SOLUTION**

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Of course water is the very best solvent :-)

$\text{Na}_2\text{S}_2\text{O}_3$  solubility in water is about 70 g / 100 ml at room temp. and about 230 g / 100 ml at 100°C.

### **What is the best solvent for sodium thiosulfate**

Under some circumstances it is possible to prepare a solution which behaves anomalously and contains more solute

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than a saturated

Saturated Solution

solution. Such a solution is said to be supersaturated. A good example of

supersaturation is

provided by  $\text{Na}_2\text{S}_2\text{O}_3$ ,

sodium thiosulfate,

whose solubility at

$25^\circ\text{C}$  is 50 g  $\text{Na}_2\text{S}_2\text{O}_3$

per 100 g  $\text{H}_2\text{O}$ .

**10.16: Saturated and Supersaturated Solutions - Chemistry ...**

sodium thiosulfate



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Crystallization of supersaturated sodium

thiosulfate solution is initiated by the addition of the seed crystal. Crystallization

of supersaturated sodium thiosulfate -

YouTube The solid crystals in the

hydrated crystals will dissolve into the bath, forming a

supersaturated solution. When the solution for sodium

thiosulfate is gradually

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cooled the super-saturated solution should remain liquid.

## **Sodium Thiosulfate Supersaturated Solution**

Read Free Sodium Thiosulfate Saturated Solution solution is initiated by the addition of the seed crystal. Crystallization of supersaturated sodium thiosulfate Solution Preparation (work with a partner)

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Weigh the amount of sodium thiosulfate necessary to make 500 ml of 0.025 M solution (this is the answer to your pre-lab question 3) using a top

### **Sodium Thiosulfate Supersaturated Solution Lab**

Crystallization of supersaturated sodium thiosulfate solution is initiated by the addition of the seed crystal.

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**Crystallization of supersaturated sodium thiosulfate - YouTube**

Be aware of the concentration units in the figures: wt%: Mass of solute/total mass of solution\*100% mol/kg: Molality = moles of solute/kg of water mol/liter: Molarity = moles of solute/liter of solution Values are tabulated below the figures. See also

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density of aqueous solutions of inorganic chlorides, inorganic potassium salts, some other inorganic substances, organic acids and organic ...

**Density of aqueous solutions of inorganic sodium salts**

Sodium thiosulfate solution 0.05 M  
Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub> (0.05 N),  
Titripur®; find  
Supelco-1.60311

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MSDS, related peer-reviewed papers, technical documents, similar products & more at Sigma-Aldrich.

**Sodium thiosulfate solution 0.05 M**

**Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub> (0.05 N ...**

Molar mass of Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub>  
.5H<sub>2</sub>O = 2\*23 + 32\*2 + 3\*  
16 + 5\*18 = 280

Equivalent

mass = molar

mass / total positive

valency = 280 / 2 = 140

So 140 g of sodium

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thiosulphate dissolved  
in 1 litre will form 1N  
solution 14 of it  
dissolved in 1 litre of  
water will give .1N  
solution

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