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Stoichiometry SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1.

_____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products.

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Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according ...

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Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio SECTION 1 Defining Stoichiometry 368 Chapter 11 • Stoichiometry Program: Chemistry Component: SE PDF Vendor: Symmetry National Chapter 11 Charles D. Winters/Photo Researchers 0368_0372_C11_S1_896405.indd 368 2/10/11 11:24 AM

CHAPTER 11 Stoichiometry

Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided.

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1. ____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

CHAPTER 9 REVIEW

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Stoichiometry. SECTION 2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles of O_2 form if 3.0 mol of KClO_3 are totally consumed? 2. Given the following equation: $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$

CHAPTER 9 REVIEW

CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$

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1 mol Al 2O_3 or 1 mol Al 2O_3 ____ 101.96 g Al 2O_3 26.98 g Al ____ 1 mol Al or ____ 1 mol Al 26.98 g Al 32.00 g O_2 ____ 1 mol O_2 or 1 mol O_2 ____ 32.00 g O_2 To find the number of grams of aluminum equivalent to 26.0 mol of aluminum, the calculation would be as follows. 26.0 mol Al \times 26.98 g Al ____ 1 mol Al = 701 g Al Stoichiometry 291 SECTION 1 ...

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Modern Chemistry Chapter 9 Homework 9 1 Answers

1. Write the definition of reaction stoichiometry in your own words. Introduction to Stoichiometry SECTION 9.1 amount of given substance (mol) convert into amount of unknown substance (mol) Ratios of substances in chemical reactions can be used as conversion factors. Reaction stoichiometry problems can be approached by looking

SECTION 9.1 Introduction to Stoichiometry

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