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Stoichiometry problems of this type are called either mass-volume or volume-mass problems. Because both types of problems involve a conversion from either moles of gas to volume or vice-versa, we can use the molar volume of 22.4 L/mol provided that the conditions for the reaction are STP.

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Simple stoichiometry only (one given, one wanted) Limiting reagents only (two given reactants, one wanted product) Mix & match (both simple stoichiometry and limiting reagent problems) Units to use (select at least one): Grams Moles Particles (e.g. atoms/molecules/formula units) Chemical formulas or names:

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Chemistry Chp. 9 Review | CourseNotes

Summary notes, flashcards and past exam questions by topic for CIE IGCSE Chemistry Topic 4 - Stoichiometry

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Practice Test Ch3 Stoichiometry (page 3 of 3) 1. d It might be easiest to balance the equation with mostly whole numbers: 2

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$\text{NH}_3 + \frac{7}{2}\text{O}_2 \rightarrow 2\text{NO}_2 + 3\text{H}_2\text{O}$. The question asks about the amount of oxygen reacting with ONE mole of ammonia, thus cut the $\frac{7}{2}$ (3.5) of oxygen in half to 1.75. c Balance: CH

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