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## **Chemistry Answer Key Heat Phase Changes**

Chemistry Answer Key  
Heat Phase Unit 3 Test:  
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1) Solve for the heat  
required to increase  
the water temperature  
from 33.0 oC to 100.0  
oC.

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Graphite is the most stable phase of carbon at normal conditions. On the phase diagram, label the graphite phase. If graphite at normal conditions is heated to 2500 K while the pressure is increased to  $10^{10}$  Pa, it is converted into diamond. Label the diamond phase.

## 8.1: Heating Curves and Phase ... - Chemistry

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3. Boiling. Heat is being added to the water to get it from the liquid state to the gas state. 4. Freezing. Heat is exiting the system in order to go from liquid to solid. Another way to look at it is to consider the opposite process of melting.

## **7.3: Phase Changes - Chemistry LibreTexts**

that they understand



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the nature to the  
science of chemistry.

5. Standardized  
Test/State Test

Connections:

Summarizing main idea  
Lesson Two and Three -  
Heat and Matter

Vocabulary Research

A. The learner will  
research the meaning  
of vocabulary terms  
related to the subject

1. Lesson Content: a.  
Phase Changes: Heat  
and Matter 2. Concept

Objective: a.

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## **Chemistry: Phase Changes: Heat and Matter**

524.8 kJ Start with  
Specific Heat because  
the water is not going  
through a phase  
change.  $q = mC. p. \Delta T$   
 $q = (200g)(4.18$   
 $J/g^{\circ}C)(75^{\circ}C) q = ? q =$   
 $62700 J = 62.7 kJ. m =$   
 $200 g C. p. = 4.18$   
 $J/g^{\circ}C \Delta T = 100^{\circ}C - 25^{\circ}C$   
 $= 75^{\circ}C. Next, the$   
water boils so you use  
Heat of Vaporization.

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Phase changes. Phase  
changes involve  
absorption or emission  
of heat. So the transfer  
of heat into a system  
can change the  
temperature of the

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system, or the phase,  
or both. KJF §17.6.

Temperature change.

(no phase change) The  
amount of heat

Q needed to increase  
the temperature of a  
mass  $m$  from  $T_1$  to  $T_2$ .

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start taking biology for this reason. The materials in this section were developed for a regular-level high school chemistry course. I have a test bank and answer key set written specifically for the resources on this website.

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When a solid is converted to a liquid, heat is absorbed. The average kinetic energy of the system changes while all of a solid is converted to a liquid. When heat is applied to a solid, the molecular motion decreases as the temperature increases. Temperature remains constant while all of a solid is converted to a liquid.

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Phase Change ...

## **Heat And Phase Changes Answer Key**

Purpose: This  
Powerpoint provides a  
basic outline of the  
important concepts  
from this chapter,  
starting with  
temperature and phase  
changes, then moving  
through the three gas

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laws: Boyle's, Amonton's, and Charles'. Real-life examples of the different laws are provided when possible. The chapter ends with the ideal gas law. Essential Concepts: Temperature, kinetic energy, Celsius, Fahrenheit ...

## **Gas Laws, Specific Heat, and Phase Changes Lessons**

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Calculating Energy for  
Changes of Phase To  
Calculate Heat: Heat to  
Change the

Temperature:  $Q = m \cdot$

$T \cdot s$  Heat to Melt:  $Q =$

$m \cdot H_{\text{fus}}$  Heat to Boil:

$Q = m \cdot H_{\text{vap}}$  For

Water:  $H_{\text{fus}} = 334 \text{ J/g}$

$H_{\text{vap}} = 2260 \text{ J/g}$

Specific Heat: ice =

$2.06 \text{ J/g } ^\circ\text{C}$  water =

$4.184 \text{ J/g } ^\circ\text{C}$  steam =

$2.03 \text{ J/g } ^\circ\text{C}$

## Calculating Energy

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The simulation includes a teacher guide, student activity sheet, and an answer key. Students use a team-based approach to solve the problem of upscaling a chemical process from lab scale to production scale for a hypothetical reaction with the Heat Flow Process Engineering Optimization activity.



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Classroom

**Resources | Phase  
Changes and Heat  
Transfer Unit ...**

Correct answers: 3  
question: A sample of ice absorbs 15.6kJ of heat as it undergoes a reversible phase transition to form liquid water at 0°C. What is the entropy change for this process in units of JK? Report your answer to three significant figures. Use

-273.15°C for

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Changes

absolute zero.

**A sample of ice  
absorbs 15.6kJ of  
heat as it undergoes  
a ...**

Unit 3: Phases of  
Matter-key Regents  
chemistry 14- Z15 Mr.  
Murdoch Page 15 of 63  
Website upload 2014  
Phase Change  
Diagrams  
(Heating/Cooling  
Curve) Page 14: When  
applying CONSTANT  
HEAT to a solid (point

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A), the temperature of the solid increases (segment AB) until the melting point (point B) is reached, for example of water ice at  $0^{\circ}$  or 273 K.

## **Unit 3: Phases of Matter-key Regents chemistry 14 Mr ...**

The Heat and Phase Changes Worksheet is one of the most useful and informative worksheets used in scientific research. It

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was introduced in the early nineteen hundreds and is the most commonly used type of formula by which scientists record their data and work out their theories.

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